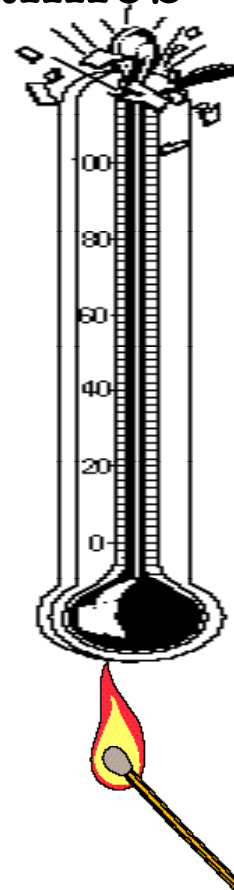
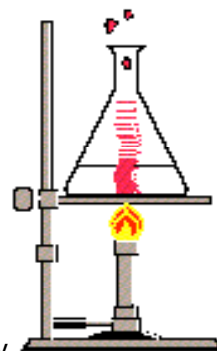


<u>Due Date</u>	<u>Assignment</u>
Thur 1/6	___ Read <u>Calorimetry Lab</u>
Fri 1/7	___ Do <u>WS 7.1</u> ___ Do <u>Calorimetry Lab ?'s</u>
Mon 1/10	___ Do <u>WS 7.2</u>
Tue 1/11	___ Do <u>WS 7.3 #1-12</u>
Wed 1/12	___ Do <u>WS 7.3 #13-17</u> ___ Read <u>Hess's Law Lab</u>
Thur 1/13	___ Do <u>Hess's Law Lab ?'s</u> ___ Read <u>Heats of Combustion Lab</u>
Fri 1/14	___ Do <u>Heats of Combustion Lab ?'s</u>
Tue 1/18	___ Do Lab Write-Up (mass using calorimetry,

+ Packet 7: Thermodynamics

$$q = m c \Delta T$$



--- *the return of Dr. Don Showalter* ---

Thur 1/20 ___ Do WS 7.4

... Mustard Day ...

Fri 1/21 ___ Do WS 7.5

___ Do WS 7.6

*** Quiz Today ***

$$\Delta G = \Delta H - T\Delta S$$

free energy

___ Come to class with packets ready to be turned in at beginning of class, with the above underlined assignments in order, in your folder, with this page as the cover page.

- For 1/2 point, be sure this packet is inside a **pocket folder**. (no torn folders or 3-ring binders)
- For 1/2 point, be sure **not** to turn in material from **1st semester**

packet order:
- assignment sheet
- WS 7.1 ~ 7.6

calorimetry:
 $q_{hot} = q_{cold}$

Some things to think about...

How is it possible to add heat to something without changing its temperature?

Could a substance exist that has a **negative** specific heat?