

+WS 3.10 Review

The following random compounds need either names or formulas. Help them discover their identity!

1. Cs_3P _____
2. SiCl_4 _____
3. platinum (IV) iodate _____
4. calcium nitrite _____
5. stannous chloride _____
6. HBrO_3 _____
7. sulfur trioxide _____
8. $\text{Fe}(\text{NO}_3)_2$ _____
9. $(\text{NH}_4)_4\text{C}$ _____
10. hydrochloric acid _____
11. mercury (II) oxalate _____
12. CoCO_3 _____
13. CO _____
14. How many electrons were transferred to make AlCl_3 ? _____

Write the Lewis dot structures, geometry, and polarity for the following:

<p>15. O_2</p> <p>electronic geometry?</p> <p>molecular geometry?</p> <p>polar? ____</p>	<p>16. HF</p> <p>electronic geometry?</p> <p>molecular geometry?</p> <p>polar? ____</p>	<p>17. BCl_3</p> <p>electronic geometry?</p> <p>molecular geometry?</p> <p>polar? ____</p>
<p>18. SO_3</p> <p>electronic geometry?</p> <p>molecular geometry?</p> <p>polar? ____</p>	<p>19. NO_3^{1-}</p> <p>electronic geometry?</p> <p>molecular geometry?</p> <p>polar? ____</p>	<p>20. CH_4O</p> <p>electronic geometry?</p> <p>molecular geometry?</p> <p>polar? ____</p>

Classify the bond type as non polar, polar, or ionic by calculating ΔEN :

21. KBr $\Delta\text{EN} =$ _____ bond type = _____
22. CO $\Delta\text{EN} =$ _____ bond type = _____
23. SiI_2 $\Delta\text{EN} =$ _____ bond type = _____

24. Draw all possible resonance structures for sulphur trioxide: